

Applications: Transition Metal Complexes

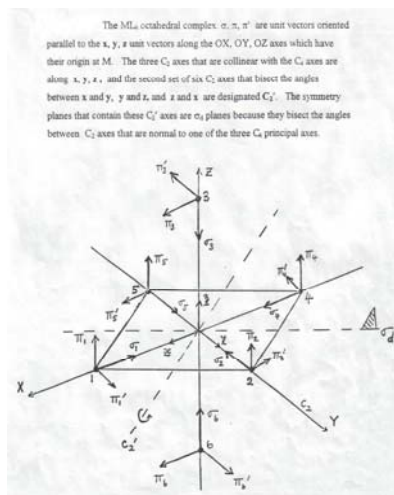
C734b

C734b transition metal complexes

1

Example:

$ML_6 O_h$ complex using coordinate system designated below:



C734b transition metal complexes

2

Metal Orbitals

nd (5)
(n+1)s (1)
(n+1) (3)

9 orbitals total

Ligand Orbitals

6 p-orbitals pointing to M (σ -orbitals)
12 p-orbitals perpendicular to σ
 $\equiv \pi$ (6 in total), π' (6 in total)

18 orbitals total

Overall orbital count = 27. This means in principle we need a 27 x 27 matrix representation!

Step 1

Examine the character table for O_h . (for the metal)

Find (although not stated) that s transforms as A_{1g}

The p-orbitals transform as T_{1u}

The d-orbitals transform as $E_g \oplus T_{2g}$

Note: the largest degeneracy is 3-fold in cubic symmetry. Therefore d-orbital set split.

The T_{2g} set (d_{xy} , d_{yz} , d_{xz}) "point" between the ligands.

The E_g set (d_{z^2} , $d_{x^2-y^2}$) "point at the ligands."

Therefore $E(T_{2g}) < E(E_g)$

These are crystal field arguments based on electrostatic repulsion between central atom electrons and ligand electrons

Step 2

Write down the characters of the representations based on ligand p-orbitals labelled σ

No need to construct matrix representations. Use "quick" method of noting how contours of the basis functions transform under symmetry operator (T)

Assign +1 to $\chi(T)$ if there is no change, -1 if the contour changes sign, and 0 if the orbital transforms to another orbital (implying an off diagonal matrix element in the matrix representation).

Characters for the irreducible representations of the point group $O_h = O \otimes C_i$.

O_h	E	$8C_3$	$6C_2'$	$6C_4$	$3C_2$	I	$\frac{8C_3}{8S_6}$	$\frac{6C_2'}{6\sigma_d}$	$\frac{6C_4}{6S_4}$	$\frac{3C_2}{3\sigma_h}$
A_{1g}	1	1	1	1	1	1	1	1	1	1
A_{2g}	1	1	-1	-1	1	1	1	-1	-1	1
E_g	2	-1	0	0	2	2	-1	0	0	2
T_{1g}	3	0	-1	1	-1	3	0	-1	1	-1
T_{2g}	3	0	1	-1	-1	3	0	1	-1	-1
A_{1u}	1	1	1	1	1	-1	-1	-1	-1	-1
A_{2u}	1	1	-1	-1	1	-1	-1	1	1	-1
E_u	2	-1	0	0	2	-2	1	0	0	-2
T_{1u}	3	0	-1	1	-1	-3	0	1	-1	1
T_{2u}	3	0	1	-1	-1	-3	0	-1	1	1

C734b transition metal complexes

5

σ -orbitals

O_h	E	$8C_3$	$6C_2'$	$6C_4$	$3C_2$	i	$8S_6$	$6\sigma_d$	$6S_4$	$3\sigma_h$
Γ_σ	6	0	0	2	2	0	0	2	0	4

Reduction yields: $\Gamma_\sigma = A_{1g} \oplus E_g \oplus T_{1u}$

C734b transition metal complexes

6

Step 3

σ -bonding. Need to use projection operators to find linear combination of ligand σ -orbitals having A_{1g} , E_g and T_{1u} symmetry

Hard work!!

Omitting normalization factors, they are:

$$\begin{aligned} \psi(A_{1g}) = & 1(\sigma_1) + 1(\sigma_2 + \sigma_3 + \sigma_5 + \sigma_6 + \sigma_3 + \sigma_5 + \sigma_2 + \sigma_6) \\ & E \qquad \qquad \qquad 8C_3 \\ & + 1(\sigma_2 + \sigma_5 + \sigma_4 + \sigma_4 + \sigma_3 + \sigma_6) + 1(\sigma_2 + \sigma_5 + \sigma_6 + \sigma_3 + \sigma_1 + \sigma_1) \\ & \qquad \qquad \qquad 6C_2' \qquad \qquad \qquad 6C_4 \\ & + 1(\sigma_4 + \sigma_4 + \sigma_1) + 1(\sigma_4) + 1(\sigma_5 + \sigma_6 + \sigma_2 + \sigma_3 + \sigma_6 + \sigma_2 + \sigma_3 + \sigma_5) \\ & \qquad \qquad \qquad 3C_2 \qquad \qquad i \qquad \qquad \qquad 8S_6 \\ & + 1(\sigma_5 + \sigma_2 + \sigma_1 + \sigma_1 + \sigma_6 + \sigma_3) + 1(\sigma_2 + \sigma_5 + \sigma_6 + \sigma_3 + \sigma_4 + \sigma_4) + 1(\sigma_1 + \sigma_1 + \sigma_4) \\ & \qquad \qquad \qquad 6\sigma_d \qquad \qquad \qquad 6S_4 \qquad \qquad \qquad 3\sigma_h \end{aligned}$$

C734b transition metal complexes

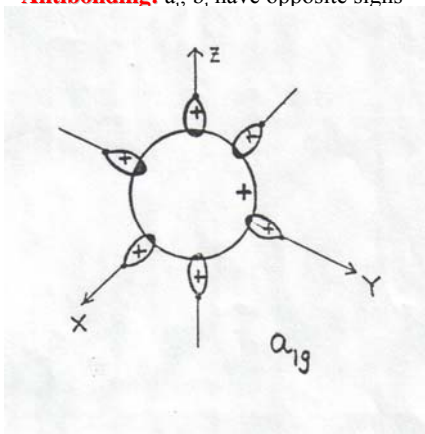
7

Therefore $\psi(A_{1g}) = \sigma_1 + \sigma_2 + \sigma_3 + \sigma_4 + \sigma_5 + \sigma_6$

$$\therefore a_{1g} = a_1[(n+1)s] + b_1(\sigma_1 + \sigma_2 + \sigma_3 + \sigma_4 + \sigma_5 + \sigma_6)$$

Bonding: a_1, b_1 have the same sign

Antibonding: a_1, b_1 have opposite signs



C734b transition metal complexes

8

Similarly:

$$\begin{aligned}\psi_1(E_g) &= 2\sigma_1 - 2(\sigma_2 + \sigma_3 + \sigma_5 + \sigma_6) + 2(\sigma_1 + 2\sigma_4) \\ &\quad + 2\sigma_4 - 2(\sigma_2 + \sigma_3 + \sigma_5 + \sigma_6) + 2(2\sigma_1 + \sigma_4)\end{aligned}$$

$$\Rightarrow \psi_1(E_g) = 2\sigma_1 - \sigma_2 - \sigma_3 + 2\sigma_4 - \sigma_5 - \sigma_6$$

Starting with σ_2 and then σ_3 gives cyclic permutations:

$$\psi_2(E_g) = 2\sigma_2 - \sigma_3 - \sigma_4 + 2\sigma_5 - \sigma_6 - \sigma_1$$

$$\psi_3(E_g) = 2\sigma_3 - \sigma_4 - \sigma_5 + 2\sigma_6 - \sigma_1 - \sigma_2$$

C734b transition metal complexes

9

Can't have 3 linearly independent basis functions for E_g

Want 2 combinations that are orthogonal (in the ZOA) and which overlap M-atom orbitals of E_g symmetry: dz^2 , dx^2-y^2

A suitable choice is:

$$\psi_1(E_g) - \psi_2(E_g) = \sigma_1 - \sigma_2 + \sigma_4 - \sigma_5 \quad (\text{xy plane})$$

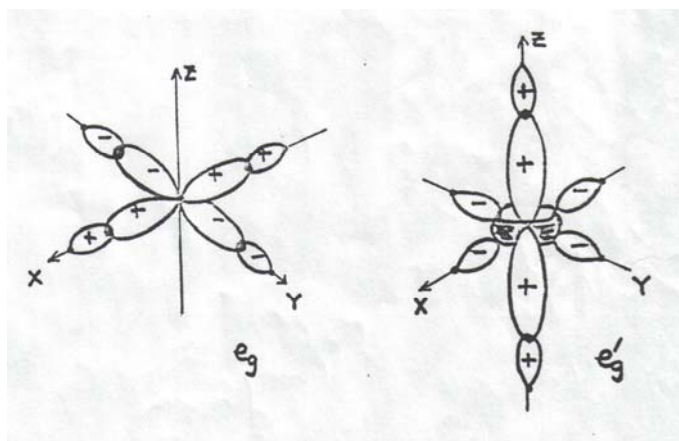
$$\psi_3(E_g) = 2\sigma_3 - \sigma_4 - \sigma_5 + 2\sigma_6 - \sigma_1 - \sigma_2 \quad (\text{z-direction})$$

C734b transition metal complexes

10

$$\therefore e_g = a_2 [nd_{x^2-y^2}] + b_2 (\sigma_1 - \sigma_2 + \sigma_4 - \sigma_5)$$

$$e'_g = a_3 [nd_{3z^2-r^2}] + b_3 (2\sigma_3 + 2\sigma_6 - \sigma_1 - \sigma_2 - \sigma_4 - \sigma_5)$$



C734b transition metal complexes

11

Similarly for the T_{1u} representation

$$\begin{aligned} \psi_1(T_{1u}) = & 3\sigma_1 - 1(\sigma_2 + \sigma_5 + \sigma_3 + \sigma_6 + 2\sigma_4) + 1(\sigma_2 + \sigma_5 + \sigma_6 + \sigma_3 + 2\sigma_1) \\ & - 1(2\sigma_4 + \sigma_1) - 3\sigma_4 + 1(\sigma_5 + \sigma_2 + 2\sigma_1 + \sigma_6 + \sigma_3) - 1(\sigma_2 + \sigma_5 + \sigma_6 + \sigma_3 + 2\sigma_4) \\ & + 1(2\sigma_1 + \sigma_4) \end{aligned}$$

$$\Rightarrow \boxed{\psi_1(T_{1u}) = \sigma_1 - \sigma_4} \quad (\text{x-direction})$$

By cyclic permutation

$$\psi_2(T_{1u}) = \sigma_2 - \sigma_5 \quad (\text{y-direction})$$

$$\psi_3(T_{1u}) = \sigma_3 - \sigma_6 \quad (\text{z-direction})$$

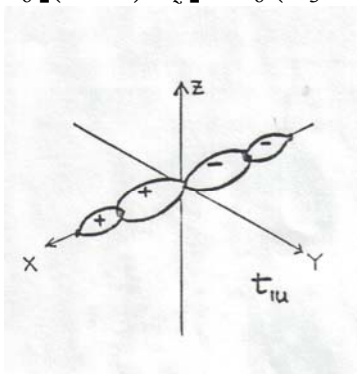
C734b transition metal complexes

12

$$\therefore t_{1u} = a_4[(n+1)p_x] + b_4(\sigma_1 - \sigma_4)$$

$$t'_{1u} = a_5[(n+1)p_y] + b_5(\sigma_2 - \sigma_5)$$

$$t''_{1u} = a_6[(n+1)p_z] + b_6(\sigma_3 - \sigma_6)$$



C734b transition metal complexes

13

Repeat procedure for the π , π' orbitals

π , π' -orbitals

O_h	E	$8C_3$	$6C_2'$	$6C_4$	$3C_2$	i	$8S_6$	$6\sigma_d$	$6S_4$	$3\sigma_h$
Γ_π	12	0	0	0	-4	0	0	0	0	$(-4+4)=0$
										$\pi \quad \pi'$

Reduction yields: $\Gamma_\pi = T_{1g} \oplus T_{2g} \oplus T_{1u} \oplus T_{2u}$

C734b transition metal complexes

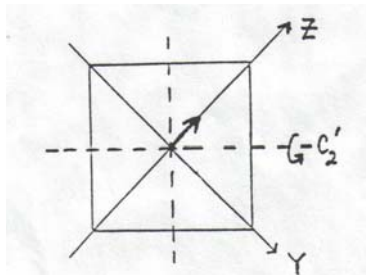
14

Next, use projection operators to find linear combinations of ligand π and π' orbitals that form bases for the T_{1u} , T_{2g} , T_{2u} , and T_{1g} IRs

Start with π_1 as our arbitrary function (**hard work!!**).

Sometimes hard to visualize.

Diagram showing for example how π_1 transforms under the C_2' operator in the yz plane



C734b transition metal complexes

15

T_{1u}

$$\begin{aligned} \psi_1(T_{1u}) = & 3(\pi_1) - 1(-\pi_2 - \pi_5 + \pi_4' - \pi_4 + \pi_3 - \pi_6) \\ & + 1(\pi_2 + \pi_5 + \pi_1' - \pi_1 + \pi_6 - \pi_5) - 1(\pi_4 - \pi_1 - \pi_4) \\ & - 3(-\pi_4) + 1(\pi_5 + \pi_2 - \pi_1' + \pi_1 - \pi_6 + \pi_3) \\ & - 1(-\pi_5 - \pi_2 - \pi_4' + \pi_4 - \pi_3 + \pi_6) + 1(-\pi_1 + \pi_4 + \pi_1) \end{aligned}$$

$$\Rightarrow \psi_1(T_{1u}) = \pi_1 + \pi_2 + \pi_4 + \pi_5 = \pi_z$$

Since O_x and O_y are equivalent in O_h symmetry

$$\psi_2(T_{1u}) = -\pi_2' + \pi_3 + \pi_5' + \pi_6 = \pi_x$$

$$\psi_3(T_{1u}) = \pi_1' + \pi_6' - \pi_4' + \pi_3' = \pi_y$$

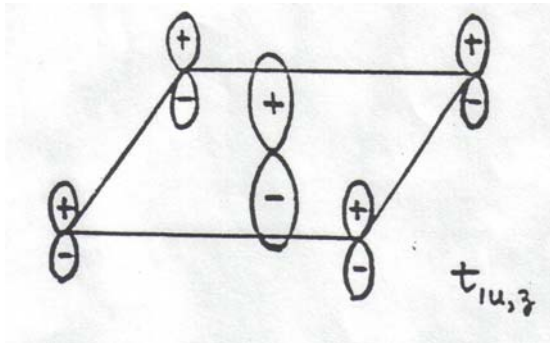
C734b transition metal complexes

16

MOs which form bases for T_{1u} are:

$$t_{1u,x} = a_7[(n+1)p_x] + b_7\pi_x$$

$$t_{1u,y} = a_8[(n+1)p_y] + b_8\pi_y$$

$$t_{1u,z} = a_9[(n+1)p_z] + b_9\pi_z$$


C734b transition metal complexes

17

T_{2g}

$$\begin{aligned} \psi_1(T_{2g}) &= 3(\pi_1) + 1(-\pi_2 - \pi_5 + \pi_3 - \pi_6) - 1(\pi_2 + \pi_5 + \pi_6 - \pi_3) \\ &\quad - 1(-\pi_1) + 3(-\pi_4) + 1(\pi_5 + \pi_2 - \pi_6 + \pi_3) \\ &\quad - 1(-\pi_5 - \pi_2 - \pi_3 + \pi_6) - 1(\pi_4) \end{aligned}$$

$$\Rightarrow \psi_1(T_{2g}) = \pi_1 - \pi_4 + \pi_3 - \pi_6 = \pi_{zx}$$

Similarly

$$\psi_2(T_{2g}) = \pi_2 - \pi_3 - \pi_5 - \pi_6 = \pi_{yx}$$

$$\psi_3(T_{2g}) = \pi_1 - \pi_2 + \pi_4 - \pi_5 = \pi_{xy}$$

C734b transition metal complexes

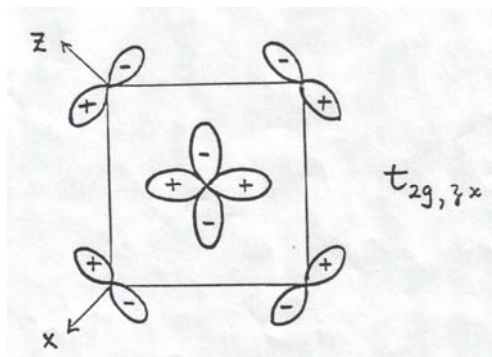
18

Therefore MOs of T_{2g} symmetry are:

$$t_{2g,xy} = a_{10}[nd_{xy}] + b_{10}\pi_{xy}$$

$$t_{2g,yz} = a_{11}[nd_{yz}] + b_{11}\pi_{yz}$$

$$t_{2g,zx} = a_{12}[nd_{zx}] + b_{12}\pi_{zx}$$



C734b transition metal complexes

19

T_{1g}

$$\begin{aligned} \psi_1(T_{1g}) &= 3(\pi_1) - 1(-\pi_2 - \pi_5 + \pi_3 - \pi_6) + 1(\pi_2 + \pi_5 + \pi_6 - \pi_3) \\ &\quad - 1(-\pi_1) + 3(-\pi_4) - 1(\pi_5 + \pi_2 - \pi_6 + \pi_3) \\ &\quad + 1(-\pi_5 - \pi_2 - \pi_3 + \pi_6) - 1(\pi_4) \end{aligned}$$

$$\Rightarrow \psi_1(T_{1g}) = \pi_1 - \pi_3 - \pi_4 + \pi_6$$

Since there is no M T_{1g} orbital means $\psi_1(T_{1g})$ is a non-bonding orbital $\equiv t_{1g,y}$

C734b transition metal complexes

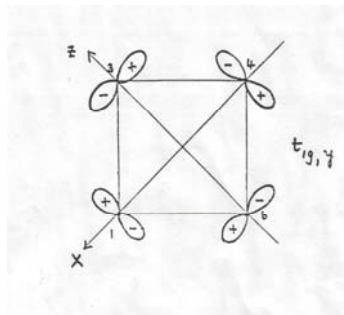
20

The 3 degenerate MO's are:

$$t_{1g,x} = b_{13}(\pi_2 + \pi_3' - \pi_5 + \pi_6')$$

$$t_{1g,y} = b_{14}(\pi_1 + \pi_3 - \pi_4 + \pi_6)$$

$$t_{1g,z} = b_{15}(\pi_1' + \pi_2' - \pi_4' + \pi_5')$$



C734b transition metal complexes

21

T_{2u}

$$\begin{aligned} \psi_1(T_{1g}) &= 3(\pi_1) + 1(-\pi_2 - \pi_5 + \pi_3 - \pi_6) - 1(\pi_2 + \pi_5 + \pi_6 - \pi_3) \\ &\quad - 1(-\pi_1) + 3(-\pi_4) - 1(\pi_5 + \pi_2 - \pi_6 + \pi_3) \\ &\quad + 1(-\pi_5 - \pi_2 - \pi_3 + \pi_6) + 1(\pi_4) \end{aligned}$$

$$\Rightarrow \psi_1(T_{2u}) = \pi_1 - \pi_2 + \pi_4 - \pi_5$$

Since there is no M T_{2u} orbital means $\psi_1(T_{2u})$ is a non-bonding orbital $\equiv t_{2u,z}$

C734b transition metal complexes

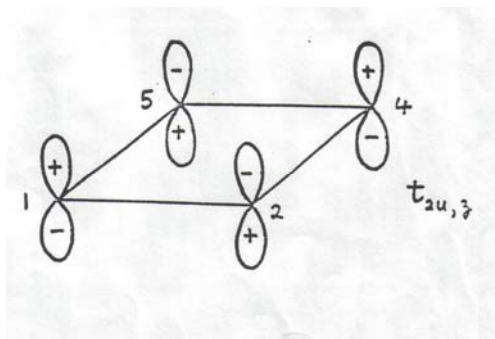
22

By symmetry the 3 degenerate MOs are:

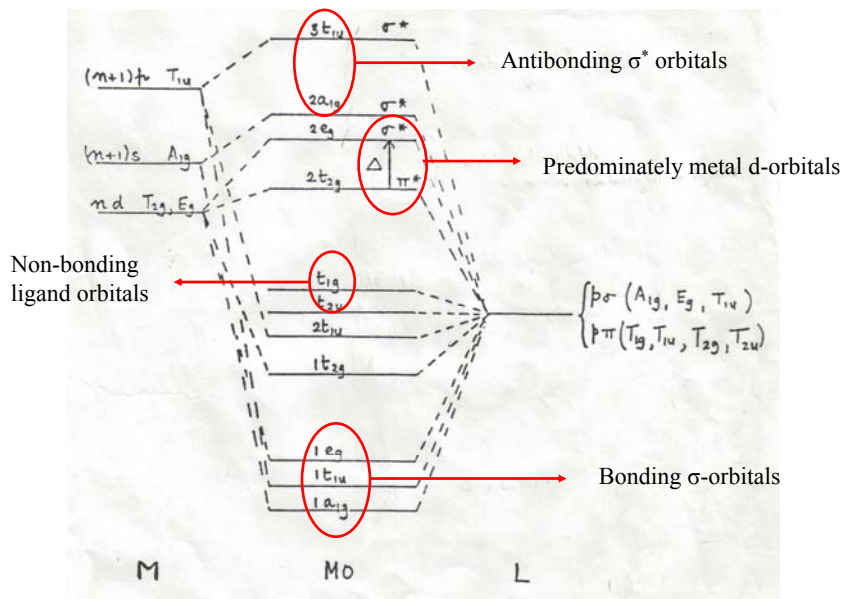
$$t_{2u,x} = b_{16}(\pi'_2 + \pi_3 - \pi'_5 + \pi_6)$$

$$t_{2u,y} = b_{17}(\pi'_1 - \pi'_6 - \pi'_4 + \pi_3)$$

$$t_{2u,z} = b_{18}(\pi_1 - \pi_2 + \pi_4 - \pi_5)$$



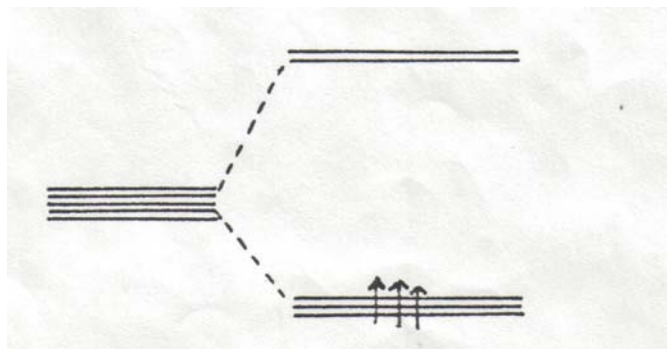
C734b transition metal complexes



C734b transition metal complexes

Crystal Field Theory picture

$(t_{2g})^3$ configuration appropriate for a $[\text{MoCl}_6]^{3-}$ complex



C734b transition metal complexes

25